

MICROCRYSTALLINE PARAFFIN, METHOD FOR PRODUCING  
MICROCRYSTALLINE PARAFFINS, AND USE OF THE  
MICROCRYSTALLINE PARAFFINS

- 5 The invention relates to a microcrystalline paraffin,  
its preparation and its use.

Conventional microcrystalline paraffin obtained from  
petroleum (also known as microwaxes) consists of a  
10 mixture of saturated hydrocarbons which are solid at  
room temperature and have a chain length distribution  
from C<sub>25</sub> to C<sub>80</sub>. The microcrystalline paraffins comprise  
n-alkanes and frequently branched isoalkanes and alkyl-  
substituted cycloalkanes (naphthenes) and also, even if  
15 generally only in small proportions, aromatics. The  
content of isoalkanes and of naphthenes ranges from 40  
to 70%, determined by the EWF Standard Test Method for  
analysis of hydrocarbon wax by gas chromatography. The  
dominance in terms of quantity of the isoalkanes (and  
20 the naphthenes) determines their microcrystalline  
structure.

The solidification range is between 50 and 100°C in  
accordance with DIN ISO 2207. The needle penetration  
25 test gives values between  $2 \times 10^{-1}$  and  $160 \times 10^{-1}$  mm in  
accordance with DIN 51579. The solidification point and  
the needle penetration are used to distinguish, among  
the microcrystalline paraffins, between plastic and  
hard microcrystalline paraffins. Soft plastic  
30 microcrystalline paraffins (known as petrolates) are  
tacky with a strongly pronounced adhesive capability  
and have solidification points from 65 to 70°C and  
penetration values from 45 to  $160 \times 10^{-1}$  mm. The oil  
contents are in the range between 1 and 15%. Plastic  
35 microcrystalline paraffins are readily shapeable and  
kneadable and have solidification points between 65 and  
80°C and penetration values from 10 to  $30 \times 10^{-1}$  mm. The  
oil contents can be up to 5%. Hard microcrystalline

paraffins are hard and tough and slightly adhesive with solidification points from 80 to 95°C and penetration values from 2 to  $15 \times 10^{-1}$  mm. The oil contents are not more than 2% (see Ullmanns Encyclopedia of Industrial Chemistry, VCH-Verlagsgesellschaft 1996).

Microcrystalline paraffins have a high molar mass and thus high boiling points. They have in the past been obtained from the residues of vacuum distillation of petroleum and from precipitates of petroleum formed during its storage (bottom residue, residual wax) using technologically very complicated and costly processes having a plurality of stages, for example deasphalting, solvent extraction, deparaffination, deoiling and refining. The deoiled microcrystalline paraffins contain sulfur, nitrogen and oxygen compounds as impurities. They are therefore not completely odorless and have a dark yellow to dark brown color. The refining which is therefore necessary is carried out as a function of the later use by bleaching (industrial applications) or by hydrorefining (applications in the food and pharmaceutical industries).

Microcrystalline paraffins are used predominantly as mixing components in paraffin and wax mixtures. The amount used is usually in the ranges up to 5%. They are employed, in particular, to increase the hardness and melting point of these mixtures and also to improve flexibility and oil binding capacity. Typical applications are, for example, the production of waxes for impregnation, coating and lamination for the packaging and textile industries, of heat sealing compounds and hot melt adhesives and also of pharmaceutical and cosmetic products, including chewing gum. Furthermore, they are used in embedding compounds and cable insulation compounds, as well as generally in plastics, and also in the candle, rubber and tire industries, as well as in cleaning, antiskid and

corrosion protection products.

It is known from WO 01/74971 that a Fischer-Tropsch product which has a broad boiling range and also  
5 contains proportions of liquid product can be isomerized under mild conditions and a wax can be isolated from the liquid hydroisomerizate by distilling off the lighter fractions. The hydrogenation temperature is given as a range from 204 to 343°C  
10 (however: 348°C in example), but the lower temperature range which is not supported by examples appears questionable in terms of its ability to be implemented. The high-boiling fractions of the starting material are blended with the wax which has been obtained in this  
15 way. A cobalt-molybdenum catalyst on aluminosilicate is reported as typical catalyst. Zeolites mentioned as being suitable are zeolite Y or ultrastable zeolite Y. The fact that an additional process step, viz. distillation, has to be employed in this proposal  
20 appears to be a disadvantage since it makes the production of the soft microcrystalline wax more expensive.

DE 69 418 388 T2 describes hydroisomerization of  
25 n-paraffins which have more than 15 carbon atoms and are solid at room temperature using a catalyst based on a metal of group VIII, in particular platinum, and a borosilicate which is isomorphous with beta-zeolite to give liquid products which are suitable for the  
30 production of lubricating oils.

DE 695 15 959 T2 describes the hydroisomerization of wax-containing starting materials to form liquid products which are suitable for the production of  
35 lubricating oils. A temperature from 270° to 360°C and a pressure from 500 to 1500 psi or from 3.44 MPa to 10.36 MPa is employed here. The hydrogen is fed into the reactor at a rate of 1000 to 10 000 SCF/bbl and the

wax is fed in at an LHSV of 0.1 to 10. The catalyst is based on a metal component on a porous, heat-resistant metal oxide support, in particular from 0.1 to 5% by weight of platinum on aluminum oxide or zeolites such as, for example, offretite, zeolite X, zeolite Y, ZSM-5, ZSM-2, etc. The starting material to be isomerized can be any wax or wax-containing material, such as, for example, slack waxes or Fischer-Tropsch wax. The isomerization product is liquid and serves as starting material for the production of lubricating oil components.

In light of this prior art, the invention addresses the object of providing a novel microcrystalline paraffin, a process for preparing it and a use of this microcrystalline paraffin.

This object is firstly and substantially achieved by the subject matter of claim 1 (product) or of claim 5 (process) or of claim 10 (use). The aim of this is that the microcrystalline paraffin which can be prepared by catalytic hydroisomerization at temperatures above 200°C, preferably 220 to 270°C, of paraffins obtained from the Fischer-Tropsch synthesis (FT paraffins) has a carbon chain length distribution in the range from C<sub>20</sub> to C<sub>105</sub>. In contrast to natural microwaxes, such a microcrystalline paraffin is free of naphthenes and aromatics. Despite isomerization, crystallinity is retained. Continuous preparation with defined properties is made possible. The preparation can even be carried out in a single process step. A product to be referred to as a microwax, having a low and high solidification point range, is provided. Continuous or batchwise catalytic hydroisomerization of T paraffins can be carried out. With regard to FT paraffins per se, reference may be made, in particular, to the information provided by A. Kühnle in Fette. Seifen. Anstrichmittel, volume 84, pages 156 ff. "Fischer-

Tropsch Wachse Synthese, Struktur, Eigenschaften und Anwendungen". Stated briefly, FT paraffins are paraffins which have been prepared by the Fischer-Tropsch process in a known manner from synthesis gas  
5 (CO and H<sub>2</sub>) in the presence of a catalyst at elevated temperature. They represent the highest-boiling fraction of the hydrocarbon mixture. The products formed are basically long-chain alkanes which have little branching and are free of naphthenes and  
10 aromatics and of oxygen compounds and sulfur compounds.

Such FT paraffins having a high proportion of n-paraffins and a carbon chain length in the range from C<sub>20</sub> to C<sub>105</sub> are converted by the process described here  
15 into high-melting, microcrystalline paraffins having a high proportion of isoparaffins.

According to the process aspect of the invention, the microcrystalline paraffin can be prepared by catalytic  
20 isomerization as follows:

- A. Use of FT paraffin as starting material
- a) having a carbon chain length in the range from C<sub>20</sub> to C<sub>105</sub>,
  - 25 b) preferably having a solidification point in the range from 70 to 105°C, in particular about 70, 80, 95 or 105°C, in accordance with DIN ISO 2207,
  - c) a penetration at 25°C from 1 to 15,
  - 30 d) a ratio of isoalkanes to n-alkanes from 1:5 to 1:11,
- B. Use of a catalyst, preferably in the form of extrudates, spheres, pellets, granules or powders,  
35 suitably based on
- a) 0.1 to 2.0% by mass, in particular 0.4 to 1.0% by mass, based on the catalyst fired at 800°C, of hydrogenating metal of transition group

eight, in particular platinum and

- b) a support material comprising a zeolite of the beta type in an amount up to 60-95% by mass, based on the combination of all components fired at 800°C,

5

- C. Use of a process temperature above 200°C, in particular from 220 to 270°C,

- 10 D. Use of a pressure from 0.5 to 20.0 MPa, preferably from 2 to 18 MPa, in particular from about 3 to 8 MPa, in the presence of hydrogen and a ratio of hydrogen to FT paraffin from 100:1 to 2000:1, in particular about 250:1 to 600:1 standard m<sup>3</sup>/m<sup>3</sup>,

15

- E. A catalyst loading with FT paraffin in the range from 0.1 to 2.0 v/v.h (volume of FT paraffin per volume of catalyst over a period of one hour), in particular 0.2 to 0.8 v/v.h.

20

In general, the catalyst fills the reactor volume virtually completely, so that it is also possible to speak of the reactor volume instead of the catalyst volume.

25

- The yield of the in each case solid hydroisomerizates is > 90% by mass, in many cases between 90 and 96% by mass, based on the respective FT paraffin used. With regard to low-melting alkanes, the hydroisomerizates obtained still contain up to 5% (in general from 2 to 3%) of alkanes having carbon chain lengths of ≤C<sub>22</sub>. These alkanes can be separated off without problems by stripping with steam under reduced pressure.

30

- 35 A catalyst which can convert solid Fischer-Tropsch paraffin selectively into microcrystalline paraffins in a single process step is used.

The catalyst comprises a combination of

- 5           - 60 to 95% by mass of zeolite of the beta type,  
          based on the combination of all components  
          fired at 800°C,
- 10          - from 5 to 39.8% by mass of a high-surface-area  
          gamma-aluminum oxide, calculated as  $\text{Al}_2\text{O}_3$  and  
          based on the combination of all components  
          fired at 800°,
- 15          - and one or more metals of transition group 8 of  
          the Periodic Table of the Elements (PTE), in  
          particular platinum, in amounts from 0.2 to  
          2.0% by mass, based on the combination of all  
          components fired at 800°C,

20       and preferably consists of this, with the metal or  
          plurality of metals of transition group 8 of the PTE,  
          in particular platinum, being attached to the aluminum  
          oxide.

25       A preferred catalyst composition for preparing  
          microcrystalline paraffins from Fischer-Tropsch  
          paraffins, in particular one having from about 20 to  
          105 carbon atoms, comprises

- 30           - from 75 to 90% by mass of zeolite of the beta  
          type, based on the combination of all  
          components fired at 800°C,
- 35          - 10 to 25% by mass of a high-surface-area gamma-  
          aluminum oxide, calculated as  $\text{Al}_2\text{O}_3$  and based on  
          the combination of all components fired at  
          800°C,
- and one or more metals of transition group 8 of  
          the Periodic Table of the Elements (PTE), in

particular platinum, in amounts from 0.4 to 1.0% by mass, based on the combination of all components fired at 800°C,

5 with the metal or the plurality of metals of transition group 8 of the Periodic Table of the PTE, in particular platinum, being attached to the aluminum oxide. The catalyst preferably consists of these abovementioned components in the abovementioned mass ratios.

10

The high-surface-area gamma-aluminum oxide has a specific surface area of 15-350 m<sup>2</sup>/g, based on  $\gamma$ -Al<sub>2</sub>O<sub>3</sub>.

15 The combination of all catalyst components fired at 800°C is free of water and ammonium.

The catalyst is preferably produced by mixing 60 to 95% by mass, in particular 75 to 90% by mass, based on the combination of all components fired at 800°C, of a preferably pulverulent zeolite of the beta type (BEA type as described by W.M. Einer, H.H. Olson & Ch. Bärlocher: Atlas of zeolite structure types, fourth ed., Elsevier London, Boston, Singapore, Sydney, Toronto, Wellington 1996),

25

- having an SiO<sub>2</sub>:Al<sub>2</sub>O<sub>3</sub> molar ratio from 19.3:1 to 100:1

30

- and a residue alkali content of not more than 0.05% by mass (based on zeolite calcined at 800°C)

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with 5 to 39.8% by mass, in particular 10 to 25% by mass, based on the combination of all components fired at 800°C, of a preferably pulverulent gamma-aluminum oxide precursor, in particular an aluminum hydroxide, preferably boehmite or pseudoboehmite, kneading the mixture with addition of water and acid as peptizing



agent, extruding the mixture, firing the extrudates at temperatures from 80°C to 200°C, in particular from 100°C to 200°C, and impregnating the shaped bodies obtained with a compound of a metal or a plurality of  
5 metals of transition group 8 of the PTE, in particular platinum, in which the noble metal is present in anionic form and subsequently thermally after-treating, in particular drying and firing, the impregnated extrudates in air, so that 0.2 to 2% by mass of noble  
10 metal, based on the combination fired at 800°C, are present, and reducing the metal or metals of transition group 8 of the PTE, in particular platinum, to the metal by means of flowing hydrogen at elevated temperature.

15 This gives shaped catalyst bodies which can be employed in a heterogeneous process, in which the catalyst is preferably used as a fixed bed and the liquefied wax together with hydrogen is passed over it at  
20 temperatures of preferably from 200 to 270°C in trickle phase mode. This catalyst combination achieves an activity in the isomerization which is sufficiently high for a Fischer-Tropsch paraffin which is solid at normal ambient temperature to be able to be used  
25 directly and a microcrystalline wax to be obtained in a single step. The properties of the microcrystalline wax can even be varied to some extent by choice of suitable reaction parameters.

30 Beta-zeolite is a commercially available product. It is preferably used according to the invention as crystalline aluminosilicate powder having the composition  $\text{Na}_n[\text{Al}_n\text{Si}_{64-n}\text{O}_{128}]$  with  $n < 7$ . In place of aluminum, boron or gallium can also be inserted  
35 isomorphously into the three-dimensional silicate structure. As a result of its high  $\text{SiO}_2$  content, it can also be exposed to an acid medium without losing its crystalline structure, although part of the aluminum

tetrahedra can be removed from the crystalline lattice. The beta-zeolite is preferably used as a fine powder having a particle size of, in particular, 0.5 to about 200  $\mu\text{m}$ , measured by means of laser particle size analyzer. The zeolite has pores having diameters from about 0.5 to 0.8 nm. The 12-membered ring openings in the structure have a width of 0.55 nm in the [001] direction and a width of 0.64 or 0.76 nm in the [100] direction of the crystal lattice. As a result of these dimensions of the openings, the long, straight-chain paraffins are obviously able, at least to some extent, to enter into the internal structure of the zeolite with its acidic sites.

For the catalyst to be particularly effective in the isomerization, the alkali cations still present after the synthesis can be replaced as quantitatively as possible by protons. The replacement of the alkali cations by protons is carried out by methods known per se, for example by exchange with water-soluble ammonium salts and subsequent calcination at 500°C. The introduction of protons can also be carried out directly using dilute acids. After the calcination, the zeolites are in the Brönstedt or or Lewis acid form (acidic sites) which is active for carbonium ion reactions.

In the production of the catalyst, the zeolite is, in a preferred embodiment, in particular used as powder together with a  $\gamma$ -aluminum oxide Al<sub>2</sub>O<sub>3</sub> which serves simultaneously as binder for the zeolite and as support for a hydrogenation metal component, or comprises these in substantial proportions. The two powders are combined with one another and, simultaneously or afterward, dilute acid, for example mineral acid, preferably nitric acid, or organic acid, e.g. formic or acetic acid, as peptizing agent and sufficient water for a shapeable, plasticized mass to be formed on

intensive working of the mass by kneading are added. To increase the plasticity, small amounts of up to about 5% by mass, based on the powder substances, of plasticizers, in particular organic auxiliaries, e.g. water-soluble cellulose ethers, are added in a preferred embodiment. This mass is extruded through nozzles, for example by means of a screw extruder, as a result of which shaped bodies in extrudate form having a chosen diameter and profile are formed. The extrudates are subsequently dried at temperatures from 80°C to 200°C, in particular from 100°C to 200°C, if desired broken to a particular length and treated thermally, in particular calcined, at temperatures from about 400°C to 600°C, in a further step, so that all or virtually all organic constituents, water and any nitrate and ammonium ions are driven off from the shaped bodies.

On calcination at temperatures above about 350°C, the aluminum oxide precursor is converted into gamma-aluminum oxide which has a specific surface area of 150 to 350 m<sup>2</sup>/g, based on Al<sub>2</sub>O<sub>3</sub>, and a pore volume from 0.3 to about 1.0 cm<sup>3</sup>/g, based on Al<sub>2</sub>O<sub>3</sub>. The pores of the aluminum oxide preferably have diameters from 3 to 59 nm, making the aluminum oxide capable of taking up large molecules and transporting them to the zeolite crystals.

The calcined shaped bodies are impregnated with a solution containing the compound of the metal or metals of transition group 8 of the PTE, in particular of platinum. Particularly useful compounds for this purpose are H<sub>2</sub>[PtCl<sub>6</sub>] and H<sub>2</sub>[PdCl<sub>4</sub>]. However, other suitable compounds in which the noble metals are present in anionic form can also be used. The compounds of the noble metals are, in a preferred embodiment, advantageously used in aqueous solution. The concentration of the noble metals in the solution is

advantageously set so that the desired final concentration of these in the catalyst is established after uptake of the solution so as to simply fill the pores of the shaped bodies with the solution.

5

After impregnation of the shaped bodies with the solution of the noble metal compounds, the shaped bodies are preferably dried in an apparatus in order to remove water. The shaped bodies are subsequently fired in a stream of dry air, with the volatile compounds becoming liberated being carried away in the offgas. Nitrous gases formed may have to be destroyed.

The noble metals are then present as finely dispersed metalloxy, in particular platinum-oxy, compounds, while the zeolite crystals themselves do not contain any hydrogenation metal component. Before being used, the catalyst is reduced in a hydrogen-containing gas stream, in particular heated to temperatures from 100 to 480°C, to deposit the noble metal in finely divided metallic form on the aluminum oxide. The metal agglomerates are advantageously and in a preferred embodiment of the invention then present in such a form that at least 30% and at most about 70% of all metal atoms are capable of adsorbing a CO molecule.

The metal components function as hydrogenation-active constituents of the catalyst which are able to activate the long-chain paraffins to form carbonium ions. The latter react at the acidic sites in the catalyst so as to displace CH<sub>3</sub> groups on the long chains. After the transformation, [lacuna] primarily in the 2-, 3-, 4- and/or 5-[lacuna] emerge from the zeolitic pore openings.

35

Beta-zeolite is a commercially available product. It is preferably used according to the invention as crystalline aluminosilicate powder having the

composition  $\text{Na}_n[\text{Al}_n\text{Si}_{64-n}\text{O}_{128}]$  with  $n < 7$ . In place of aluminum, boron or gallium can also be inserted isomorphously into the three-dimensional silicate structure. As a result of its high  $\text{SiO}_2$  content, it can also be exposed to an acid medium without losing its crystalline structure, although part of the aluminum tetrahedra can be removed from the crystalline lattice. The beta-zeolite is preferably used as a fine powder having a particle size of, in particular, 0.5 to about 200  $\mu\text{m}$ , measured by means of laser particle size analyzer. The zeolite has pores having diameters from about 0.5 to 0.8 nm. The 12-membered ring openings in the structure have a width of 0.55 nm in the [001] direction and a width of 0.64 or 0.76 nm in the [100] direction of the crystal lattice. As a result of these dimensions of the openings, the long, straight-chain paraffins are obviously able, at least to some extent, to enter into the internal structure of the zeolite with its acidic sites.

For the catalyst to be particularly effective in the isomerization, the alkali cations still present after the synthesis can be replaced as quantitatively as possible by protons. The replacement of the alkali cations by protons is carried out by methods known per se, for example by exchange with water-soluble ammonium salts and subsequent calcination at  $500^\circ\text{C}$ . The introduction of protons can also be carried out directly using dilute acids. After the calcination, the zeolites are in the Brönstedt or or Lewis acid form (acidic sites) which is active for carbonium ion reactions.

In the production of the catalyst, the zeolite is, in a preferred embodiment, in particular used as powder together with a  $\gamma$ -aluminum oxide Al00H which serves simultaneously as binder for the zeolite and as support for a hydrogenation metal component, or comprises these

- in substantial proportions. The two powders are combined with one another and, simultaneously or afterward, dilute acid, for example mineral acid, preferably nitric acid, or organic acid, e.g. formic or acetic acid, as peptizing agent and sufficient water for a shapeable, plasticized mass to be formed on intensive working of the mass by kneading are added. To increase the plasticity, small amounts of up to about 5% by mass, based on the powder substances, of plasticizers, in particular organic auxiliaries, e.g. water-soluble cellulose ethers, are added in a preferred embodiment. This mass is extruded through nozzles, for example by means of a screw extruder, as a result of which shaped bodies in extrudate form having a chosen diameter and profile are formed. The extrudates are subsequently dried at temperatures from 80°C to 200°C, in particular from 100°C to 200°C, if desired broken to a particular length and treated thermally, in particular calcined, at temperatures from about 400°C to 600°C, in a further step, so that all or virtually all organic constituents, water and any nitrate and ammonium ions are driven off from the shaped bodies.
- On calcination at temperatures above about 350°C, the aluminum oxide precursor is converted into gamma-aluminum oxide which has a specific surface area from 150 to 350 m<sup>2</sup>/g, based on Al<sub>2</sub>O<sub>3</sub>, and a pore volume from 0.3 to about 1.0 cm<sup>3</sup>/g, based on Al<sub>2</sub>O<sub>3</sub>. The pores of the aluminum oxide preferably have diameters from 3 to 59 nm, making the aluminum oxide capable of taking up large molecules and transporting them to the zeolite crystals.
- The calcined shaped bodies are impregnated with a solution containing the compound of the metal or metals of transition group 8 of the PTE, in particular of platinum. Particularly useful compounds for this

purpose are  $H_2[PtCl_6]$  and  $H_2[PdCl_4]$ . However, other suitable compounds in which the noble metals are present in anionic form can also be used. The compound of the noble metals are, in a preferred embodiment, 5 advantageously used in aqueous solution. The concentration of the noble metals in the solution is advantageously set so that the desired final concentration of these in the catalyst is established after uptake of the solution so as to simply fill the 10 pores of the shaped bodies with the solution.

After impregnation of the shaped bodies with the solution of the noble metal compounds, the shaped bodies are preferably dried in an apparatus in order to 15 remove water. The shaped bodies are subsequently fired in a stream of dry air, with the volatile compounds becoming liberated being carried away in the offgas. Nitrous gases formed may have to be destroyed.

20 The noble metals are then present as finely dispersed metalloxy, in particular platinum-oxy, compounds, while the zeolite crystals themselves do not contain any hydrogenation metal component. Before being used, the catalyst is reduced in a hydrogen-containing gas 25 stream, in particular heated to temperatures from 100 to 480°C, to deposit the noble metal in finely divided metallic form on the aluminum oxide. The metal agglomerates are advantageously and in a preferred embodiment of the invention then present in such a form 30 that at least 30% and at most about 70% of all metal atoms are capable of adsorbing a CO molecule.

The metal components function as hydrogenation-active constituents of the catalyst which are able to activate 35 the long-chain paraffins to form carbonium ions. The latter react at the acidic sites in the catalyst so as to displace  $CH_3$  groups on the long chains. After the transformation, paraffins which are singly methyl group

branched primarily in the 2-, 3-, 4- and/or 5-position on the carbon chain emerge from the zeolitic pore openings. The catalyst can, for example, be used in the form of extrudates, cylinders, granules, spheres, pellets or powder.

The catalyst is preferably used in the presence of hydrogen at an  $H_2$  partial pressure from 5 to 180 bar.

10 The catalyst is also preferably used at an  $H_2$ :feed ratio from 100:1 to 2000:1 standard  $m^3/m^3$  of feed.

The catalyst is also preferably used at a loading from 0.1 to 1 volumes of feed/volumes of catalyst and hour.

15

The catalyst is also preferably used at a temperature from 200° to 270°.

20 The catalyst can also be used in the form of small particles suspended in the feed at temperatures from preferably 200°C to 270°C and elevated pressure in the presence of hydrogen in order to convert Fischer-Tropsch paraffin into microcrystalline wax. Any light constituents occurring can be driven off by means of steam distillation (stripping).

25

30 The catalyst is advantageously installed as a fixed bed in a reactor through which the feed together with hydrogen is allowed to flow slowly at temperatures from preferably 200°C to 270°C. The catalyst can be used in a continuous, semicontinuous or batchwise mode of operation.

35 The catalyst is illustrated by means of the following example.



## Example

## Production of a catalyst

5 300 g of commercially available beta-zeolite having an  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio of 23.3 in the metal cation-free form (alkali content less than 0.05% by mass, based on zeolite calcined at 800°C) as powder having a particle size from 0.5 to about 50  $\mu\text{m}$ , 62.8 g of commercially  
10 available aluminum oxide hydroxide as fine powder and 8.4 g of water-soluble cellulose ether are intensively mixed with one another. 30 ml of dilute nitric acid containing 128 g of  $\text{HNO}_3/\text{l}$  and 350 ml of deionized water are then added and the mixture is kneaded  
15 intensively for one hour. This gives a shapeable, kneadable mass. The mass obtained in this way is extruded by means of a screw extruder through nozzles having cylindrical openings with a diameter of 1.5 mm, so that rod-shaped extrudates are formed. These are  
20 dried at 120°C for 6 hours in a drying oven. The shaped bodies are broken to a length of 3-5 mm and calcined at 550°C in a thin layer on a metal sheet for three hours in an electric muffle furnace through which air was slowly passed. Solid shaped bodies having a bulk  
25 density of 400 g/l are obtained.

The amount of water which can be taken up by the shaped bodies, which corresponds to the pore volume, is determined at room temperature (= 110% based on the  
30 mass of catalyst). A solution of 1.636 g of  $\text{H}_2\text{PcCl}_6$  in 242 ml of water is sprayed onto 220 g of the shaped bodies while the bodies are kept in motion. After the solution has been allowed to act on the shaped bodies for 10 minutes, the shaped bodies are partly dried  
35 while being kept in motion until the major part of the liquid has been vaporized and the individual shaped bodies no longer stick to one another. The impregnated shaped bodies are then dried at 120°C in air in a

drying oven. The dried shaped bodies are heated at 100°C/h to 450°C in a stream of dry air in a vertical oven and maintained at 450°C for one hour.

- 5 The shaped bodies are subsequently cooled to ambient temperature in the oven, the stream of air is replaced by pure nitrogen until the oxygen content in the out-flowing gas is less than 0.5% by volume and there is then change-over from nitrogen to hydrogen. The oven is  
10 once again heated at 100°C/h to 450°C and the catalyst is treated, i.e. reduced, at this temperature in the stream of hydrogen for three hours. The catalyst is then allowed to cool in a stream of nitrogen and can then be taken out. The resulting catalyst A according  
15 to the invention is stable in air. The platinum content is 0.8% by mass, based on the combination of all components fired at 800°C.

#### Catalytic Test

20

- The catalyst A produced above was comminuted to a particle size of 160-315 µm and 4 g of this comminuted catalyst were stirred into 180 g of a Fischer-Tropsch paraffin ("feed") at a temperature of 120°C. The  
25 mixture was placed in an autoclave. After the autoclave had been closed, it was pressurized with 50 bar of hydrogen and the mixture was heated to 250°C whilst stirring and treated further for seven hours whilst stirring. The autoclave was then cooled to 120°C again,  
30 and the product was taken from the autoclave, the catalyst was separated off and examined. The product data were compared with those of the feed (see table).

Table: Data for the feed and the hydroisomerizate

	Feed	Hydroisomerizate
Number of C atoms (>90%)	30 to 100	About 25 to 100
Solidification point in °C	97	86.5
Enthalpy of fusion $\Delta H$	>200	125
Penetration at 25°C	1-2	42
Viscosity at 120°C in cSt	~12	15.4
Isomer proportion in % by mass	~12	47
<C 22 in % by weight	0	2-3

5 The hydroisomerizate displays properties which are distinctly different from the starting material and correspond to a microcrystalline wax. The proportion of the i-paraffins is considerably increased compared to the feed.

10 The catalytic hydroisomerization of the FT paraffins is preferably carried out continuously in a flow reactor using a fixed-bed catalyst, in particular in the form of extrudates, spheres or pellets. If the reactor is, as is preferred, aligned vertically, the flow through  
15 it can be either from the top downward or from the bottom upward. However, the process can also be carried out non-continuously or semicontinuously in, for example, a stirring autoclave in a batch process in which the catalyst is present in a permeable mesh or is  
20 finely dispersed as granules or powder in the FT paraffin. The process parameters of the continuous process and the batchwise process are identical.

25 The solid microcrystalline paraffins obtained according to the invention have the following properties:

Compared to the FT paraffins used, they have somewhat lower solidification points and comprise not only n-alkanes but also a high proportion by weight of isoalkanes, in particular a higher proportion by weight of isoalkanes than of n-alkanes. The proportion of n-alkanes or isoalkanes is determined by gas chromatography. The increased degree of isomerization achieved by means of the hydroisomerization is reflected in increased penetration values, a reduced degree of crystallization and a reduced enthalpy of fusion. The products were solid, white, opaque and of a sticky consistency. They were in each case solid at ambient temperature (20°C).

The degree of crystallization is determined by X-ray diffraction analysis. It indicates the proportion of crystalline material in the product obtained as a ratio to the proportion of amorphous material. The amorphous materials lead to different diffraction of the X-rays than the crystalline materials. The needle penetration at 25°C of the products according to the invention is in the range from 20 to 100, measured in accordance with DIN 51579.

The proportion of crystalline material is, in particular, reduced as follows: while a proportion of crystalline material in the range from 60 to 75% is present in the starting material, the proportion observed in the hydroisomerizate is from 30 to 45%, in particular in the range from 35 to 40%. This proportion of crystalline material is accordingly halfway between that of microcrystalline paraffins based on petroleum and that of the starting material, viz. the FT paraffins. The proportion of crystalline material in these synthetic microparaffins therefore also closes a gap in the use properties of such products, since the physical and materials properties of such products are

generally functions of the crystallinity.

The proportions of crystalline and amorphous materials as determined by the abovementioned X-ray diffraction analysis are in each case reported in percent by mass.

The microcrystalline paraffins prepared by catalytic hydroisomerization can also be deoiled by means of a solvent. However, this does not imply that the hydroisomerization products described have a content of oily components in the normal sense. However, short-chain n-alkanes and isoalkanes are removed. When a solvent mixture of dichloroethane:toluene in a volume ratio of 95:5 and a product:solvent ratio of 1:3.6 are used at 22°C, a deoiled microcrystalline paraffin is obtained in a yield from 80 to 90% by weight, based on the hydroisomerizate used. It has the following properties:

- needle penetration: from  $1 \times 10^{-1}$  to  $7 \times 10^{-1}$  mm, in particular  $3 \times 10^{-1}$  to  $6 \times 10^{-1}$  mm, determined in accordance with DIN 51579,
- MBK-soluble material: 1.0 to 2% by weight, in particular 1.2 to 1.6% by weight, determined in accordance with ASTM D 721/87 modified so as to use MIBK,
- solidification point: about 60 to about 95°C, in particular 70 to 85°C, determined in accordance with DIN ISO 2207.

Removal of the short-chain material thus gives a very hard product, based on comparison with the grades obtained from petroleum, from the medium-hard product. The deoiled hydroisomerizate is then comparable with the hardest grades obtained from petroleum.

Owing to its properties, the microcrystalline hydroisomerizate prepared according to the invention and the corresponding deoiled microcrystalline

hydroisomerizate can be used like a microwax (see introduction). In particular, the hydroisomerizate obtained can also be oxidized. The oxidation products obtained can differ in their melting range and degree  
5 of oxidation and can be used, in particular, as base products for corrosion protection products and as products for protecting hollow spaces in and the underside of motor vehicles. They are also used in emulsions as cleaning products and release agents and  
10 as additives in printing inks and carbon paper compositions.

The acid and ester groups which are randomly distributed over the hydrocarbon chains can be reacted  
15 with inorganic or organic bases to produce formulations which are dispersible in water (emulsifiable waxes) and lead to products having very good adhesion to metals.

Further application areas are the production of  
20 impregnation, coating and lamination waxes for the packaging and textile industries, heat sealing compounds and hot melt adhesives, as blend components in candles and other wax products, in wax mixtures for drawing chalk, floor care products and automobile care  
25 products and for dental technology and pyrochemistry.

They are also constituents of light-protection waxes for the tire industry, electrical insulation materials, framework and modeling waxes for the fine casting  
30 industry and wax formulations for explosives, munitions and propellant technology.

Furthermore, such products are suitable as release agents in the pressing of chipboard, particleboard and  
35 fiberboard, in the production of ceramic parts and, owing to their retention capacity, for the production of solvent-containing cleaners, grinding and polishing pastes and as matting agents for surface coatings.

These products can also be used for formulating adhesive waxes, cheese waxes, cosmetic preparations, chewing gum bases, casting and cable insulation compositions, sprayable pesticides, vaselines, artificial firewood, lubricants and hot melt adhesives.

The synthetic microwaxes are foodsafe. Testing is carried out in accordance with FDA, § 175.250.

The invention will now be illustrated specifically by means of examples.

Example 1:

An FT paraffin having a solidification point of 97°C was catalytically isomerized by means of hydrogen at a pressure of 5 MPa (50 bar), a temperature of 270°C and a v/vh ratio of 0.3 in a flow reactor. The hydroisomerization occurring is evidenced by the data in table 1.

The catalyst comprised 0.8% by mass of platinum on  $\beta$ -zeolite and an  $\text{SiO}_2$  to  $\text{Al}_2\text{O}_3$  molar ratio of 23:1 and an aluminum oxide having a large surface area. The catalyst was in acid form. It contained less than 0.02% of alkali oxide on a dry basis.

The hydroisomerizate obtained was solid, white-opaque, odorless, slightly sticky and thus differed significantly from the hard-brittle starting material. The isoalkane content had been increased by a factor of about 5, which is reflected in the increased penetration value, the reduced proportion of crystalline material and the reduced enthalpy of fusion. The synthetic, microcrystalline paraffin prepared in this way is in terms of its properties between a plastic petroleum-based microwax and a hard

petroleum-based microwax. The hydroisomerizate is thus a paraffin which has a pronounced microcrystalline structure and whose carbon chain length distribution of 23 to 91 carbon atoms corresponds approximately to that of the starting material with 27 to 95 carbon atoms but shifted slightly to shorter chain lengths. The chain lengths were determined by gas chromatography; a corresponding gas chromatogram is attached as figure 1.

10 Example 2:

An FT paraffin having a solidification point of 71.5°C was catalytically isomerized under a hydrogen pressure of 5 MPa (50 bar) at a temperature of 250°C in an autoclave. The structural transformation which has occurred is evidenced by the data in the table.

The same catalyst as in example 1 was used.

20 The hydroisomerizate obtained was solid, white-opaque and odorless and also paste-like and slightly sticky. The isoalkane content was increased by a factor of about 5. The high degree of isomerization is reflected in the significantly increased penetration value, the reduced proportion of crystalline material and the reduced enthalpy of fusion. The microcrystalline paraffin obtained in this way has a similar but somewhat shorter carbon chain length than the FT paraffin, which is clearly indicated by the number of carbon atoms: 23 to 42 in the case of the hydroisomerizate and 25 to 48 in the case of the FT paraffin. The synthetic microcrystalline paraffin prepared in this way is comparable in terms of its properties to a soft plastic microcrystalline paraffin obtained from petroleum.

Examples 1 and 2 show that the FT paraffins, which consist predominantly of n-alkanes and have a finely



crystalline structure and a hard-brittle consistency, were converted by the process of the invention into nonfluid, paste-like or solid paraffins which have lower melting points than the starting materials. These  
5 paraffins have a high content of branched alkanes and consequently are distinguished by a microcrystalline structure with a significantly reduced degree of crystallization (compared to the starting material) and a plastic to slightly sticky consistency. The branched  
10 alkanes are predominantly methylalkanes, with the methyl groups preferably being located in the 2-, 3-, 4- or 5-position. Multiply methyl-branched alkanes were also formed to a small extent.

15 The results of examples 1 and 2, and also the properties of the starting material for comparison, are shown in the accompanying table 1.

20 A gas chromatogram corresponding to example 2 is shown in figure 2.

#### Example 3:

The starting material of example 2 was then isomerized  
25 in a flow reactor, once again using the same catalyst, to give a hydroisomerizate having somewhat different but comparable properties (cf. also table 1 in this regard) as in the autoclave experiment (example 2), at a significantly reduced process temperature of 220°C. A  
30 reactor experiment is significantly closer to industrial implementation of the hydroisomerization than is an autoclave experiment. The possibility of reducing the process temperature compared to the autoclave experiment which has thus been demonstrated  
35 allows the same to be expected in the case of example 1 when carried out industrially.

The reduction in the process temperature is also

associated with the considerable advantage that the cracking reaction which competes with such a hydroisomerization is substantially suppressed (cf. figures 1 to 3).

5

A gas chromatogram corresponding to example 3 is attached as figure 3.

10 In contrast to the microcrystalline paraffins obtained from petroleum, the fully synthetic microcrystalline paraffins prepared by means of the hydroisomerization according to the invention contain no highly branched isoalkanes, no cyclic hydrocarbons (naphthenes) and, in particular, no aromatics and sulfur compounds. They  
15 thus correspond to the highest purity requirements for microcrystalline paraffins and are thus highly suitable for use in the cosmetic and pharmaceutical industry and for packaging and preservation in the food industry.

Table 1: Properties of starting materials and reaction products

	Unit	Measurement method	Example 1		Example 2		Example 3	
			FT	Hydro-paraffin isomerizate (H8)	FT	Hydro-paraffin isomerizate (FT 70)	FT	Hydro-paraffin isomerizate (HDI 70-R)
Solidification point	°C	DIN ISO 2207	97.0	86.5	71.5	61.5	71.5	64.5
Penetration N at 25°C	0.1 mm	DIN 51579	2	42	13	98	13	79
Enthalpy of fusion	J/g	ASTM D4419	221	127	195	120	195	142
Proportion of crystalline material	% by mass	X-ray diffraction analysis	70.7	43.5	62.4	38.8	62.4	41.2
Isoalkanes	%	Gas chromatography	12	47	9	47	9	40
MIBK-soluble material	% by mass	ASTM D721-87 (modified)	0.66	14.6	0.4	23.1	0.4	15.0
Viscosity (100°C)	cSt		12 (120°C)	15.4 (120°C)	6.5	5.2	6.5	5.1